

# Magnetism in materials

Solutions - Week 03

- Let's assume that we have a 3d metal in a tetrahedral crystal.
  - Between the  $e_g$  orbitals and the  $t_{2g}$  orbitals, which one would have the highest energy due to the crystal field effect ?
  - Find the spin number  $S$  for all the 3d metal ( $3d^0$  to  $3d^{10}$ ) in the high spin and low spin case.

## Solutions

- The  $t_{2g}$  orbits are closer to ligand, so their energy are higher then the orbitals  $e_g$ .
- The spin distribution is as followed :

	$d^0$	$d^1$	$d^2$	$d^3$	$d^4$	$d^5$	$d^6$	$d^7$	$d^8$	$d^9$	$d^{10}$
	$Sc^{2+}/Ti^{4+}/V^{5+}$	$Sc^{2+}/Ti^{3+}/V^{4+}/Cr$	$Ti^{2+}/V^{3+}/Cr^{4+}$	$V^{2+}/Cr^{3+}/Mn^{4+}$	$Cr^{2+}/Mn^{3+}$	$Mn^{2+}/Fe^{3+}$	$Fe^{2+}/Co^{3+}$	$Co^{2+}$	$Ni^{2+}$	$Cu^{2+}$	$Zn^{2+}$
High Spin	$S = 0$ --- --	$S = 1/2$ --- ↑--	$S = 1$ --- ↑↑	$S = 3/2$ ↑--- ↑↑	$S = 2$ ↑↑-- ↑↑	$S = 5/2$ ↑↑↑ ↑↑	$S = 2$ ↑↑↑ ↑↑	$S = 3/2$ ↑↑↑ ↑↑	$S = 1$ ↑↑↑ ↑↑	$S = 1/2$ ↑↑↑ ↑↑	$S = 0$ ↑↑↑ ↑↑
Low Spin	$S = 0$ --- --	$S = 1/2$ --- ↑--	$S = 1$ --- ↑↑	$S = 1/2$ --- ↑↑	$S = 0$ --- ↑↑	$S = 1/2$ ↑--- ↑↑	$S = 1$ ↑↑-- ↑↑	$S = 3/2$ ↑↑↑ ↑↑	$S = 1$ ↑↑↑ ↑↑	$S = 1/2$ ↑↑↑ ↑↑	$S = 0$ ↑↑↑ ↑↑

Figure 1: Spin repartition for 3d metal in an tetrahedral crystal

- Using the spin distribution for a octahedral crystal discussed in course (for a 3d metal), determine in which situation the geometry will change due to the Jahn-Teller effect.

**Solutions** The configuration with Jahn-Teller effect are mark with 'J-T' in Fig.2

- A  $Sc^{2+}$  ion has one electron in the 3d shell. It is in an anisotropic crystal and the crystal field can be written as a potential acting on the 3d electron as  $E = AL_z^2$ . What are the lowest orbital states of the SC ion if  $A > 0$  and  $A < 0$  ?  
The spin-orbit coupling  $\lambda IS$  is much smaller than the crystal field. When this is included, what are the approximate ground states of the ion, for  $A > 0$  and  $A < 0$  ?

$d^0$	$d^1$	$d^2$	$d^3$	$d^4$	$d^5$	$d^6$	$d^7$	$d^8$	$d^9$	$d^{10}$
Sc <sup>3+</sup> /Ti <sup>4+</sup> /V <sup>5+</sup>	Sc <sup>2+</sup> /Ti <sup>3+</sup> /V <sup>4+</sup> /Cr <sup>3+</sup>	Ti <sup>2+</sup> /V <sup>3+</sup> /Cr <sup>2+</sup>	V <sup>2+</sup> /Cr <sup>3+</sup> /Mn <sup>4+</sup>	Cr <sup>2+</sup> /Mn <sup>3+</sup>	Mn <sup>2+</sup> /Fe <sup>3+</sup>	Fe <sup>2+</sup> /Co <sup>3+</sup>	Co <sup>2+</sup>	Ni <sup>3+</sup>	Cu <sup>2+</sup>	Zn <sup>2+</sup>
S = 0	S = 1/2 J-T	S = 1 J-T	S = 3/2	S = 2 J-T	S = 5/2	S = 2 J-T	S = 3/2 J-T	S = 1	S = 1/2 J-T	S = 0
— —	— —	— —	— —	↑ ↑	↑ ↑	↑ ↑	↑ ↑	↑ ↑	↑ ↑	↑ ↑
— — —	↑ — —	↑ ↑ —	↑ ↑ ↑	↑ ↑ ↑	↑ ↑ ↑	↑ ↑ ↑	↑ ↑ ↑	↑ ↑ ↑	↑ ↑ ↑	↑ ↑ ↑
				S = 1 J-T	S = 1/2 J-T	S = 0	S = 1/2 J-T			
				— —	— —	— —	↑ —			
				↑ ↑	↑ ↑	↑ ↑	↑ ↑			

Figure 2: Spin repartition for 3d metal in an octahedral crystal

**Solution** Sc<sup>2+</sup> has one 3d electron ( $L=2, S = 1/2$ ) and has 5 orbital states characterized by  $L_z = -2, -1, 0, 1, 2$ . The corresponding energy levels are then  $E(L_z) = 4A, A, 0, A, 4A$ . We will write a state with  $|L_z, S_z\rangle$ .

If  $A > 0$  the lowest orbital states are  $|0, \pm \frac{1}{2}\rangle$ .

If  $A < 0$  the lowest orbital states are  $|2, \pm \frac{1}{2}\rangle$  and  $|-2, \pm \frac{1}{2}\rangle$ .

Now the spin-orbit energy is taken into account.

If  $A > 0$  the ground states are  $|0, \pm \frac{1}{2}\rangle$  since  $\lambda L_z S_z = 0$  for both states.

If  $A < 0$  the ground states are  $|2, -\frac{1}{2}\rangle$  and  $|-2, \frac{1}{2}\rangle$ .

4. Vesta is a software used to visualize crystal structure. We will in this exercise use it to analyse Ca<sub>3</sub>Co<sub>2</sub>O<sub>6</sub> crystal. Download link, file and articles are all on moodle :

- (a) Download and install [Vesta](#).
- (b) Load the structure of Ca<sub>3</sub>Co<sub>2</sub>O<sub>6</sub>.
- (c) Determine the environment of Co ions and discuss possible arrangements of d-orbital.
- (d) Compare the results obtained by measurements of susceptibility, magnetization and neutron diffraction.[1][2]

5. Let us consider a compound ML<sub>6</sub> where six ligands (L) occupy the vertices of an octahedron, and the metal ion (M) resides at the center. We establish a global right-handed coordinate system aligned with the diagonals of the tetrahedron.

One of the normal modes of vibration associated with ML<sub>6</sub> is the *symmetric stretching*, depicted in (a) where ligands move along the green arrows. The normal coordinate for this mode is given by:

$$Q_1 = \frac{1}{\sqrt{6}}(X_2 - X_5 + Y_3 - Y_6 + Z_1 - Z_4)$$

where  $(X_i, Y_i, Z_i)$  represent deviations of the ligands from their equilibrium positions.  $Q_1$  equals zero when all ligands are in equilibrium,  $Q_1$  is positive when ligands move outward, and negative when ligands move inward.

In (b) and (c), two additional normal modes termed *asymmetric stretching* are depicted. The green arrows illustrate ligand motion for these modes. Find the normal coordinates for them.

**Solution** For (b), the normal coordinate is given by

$$Q_2 = \frac{1}{2\sqrt{3}}(-X_2 + X_5 - Y_3 + Y_6 + 2Z_1 - 2Z_4).$$

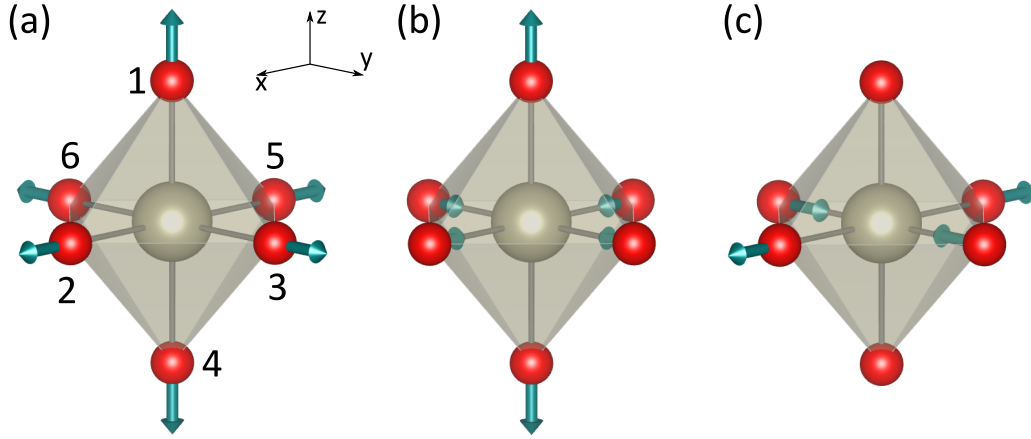


Figure 3: Normal modes of atoms in an octahedral complex  $ML_6$ .

When ligands 2, 5, 3, 6 move inward (outward), and 1, 4 move outward (inward),  $Q_2 > 0$  ( $Q_2 < 0$ ).

The normal coordinate for (c) is given by

$$Q_3 = \frac{1}{\sqrt{6}}(X_2 - X_5 - Y_3 + Y_6).$$

In this case, ligands 1, 4 do not move. When ligands 3, 6 move inward (outward), and 2, 5 move outward (inward),  $Q_3 > 0$  ( $Q_3 < 0$ ).

*Note:* Generally, electronic states can couple to multiple normal modes or their linear combinations. To visualize this phenomenon, we can utilize the data and simulations provided in the file `Jahn-Teller.ipynb`. Specifically, we aim to depict the Mexican hat energy surface depicting the coupling between electronic states and the vibrational modes  $Q_2$  and  $Q_3$ , as introduced in the previous exercise, in the case of linear Jahn–Teller effect within the harmonic approximation.

## References

- [1] A. Maignan, C. Michel, A. C. Masset, C. Martin, and B. Raveau. Single crystal study of the one dimensional  $Ca_3Co_2O_6$  compound: five stable configurations for the ising triangular lattice. *The European Physical Journal B - Condensed Matter and Complex Systems*, 15(4):657–663, Jun 2000.
- [2] B. Haubackb S. Aasland, H. Fjellvag. Magnetic properties of the one-dimensional  $Ca_3Co_2O_6$ . *Solid State Communications*, 101(3):187–192, 1997.