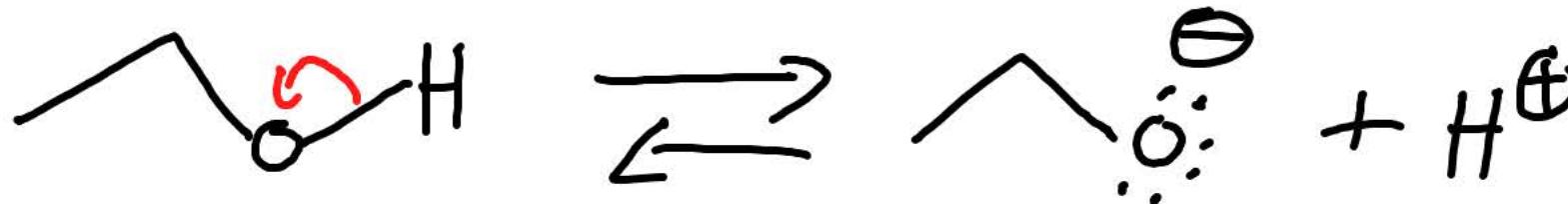
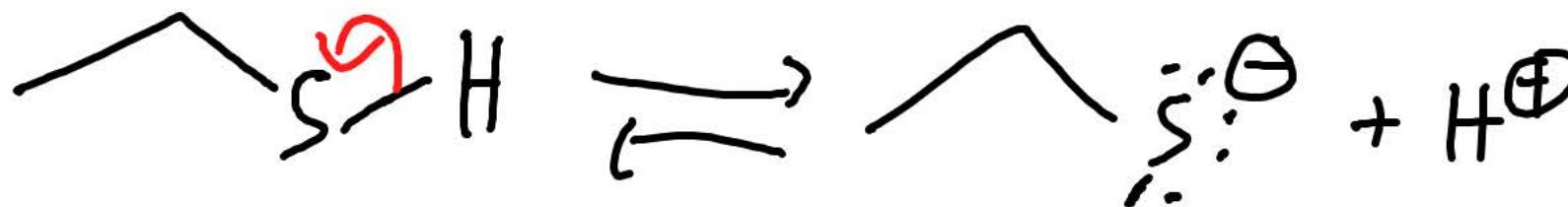


cas "inexplicable"

qui est plus acide?



pKa = 16

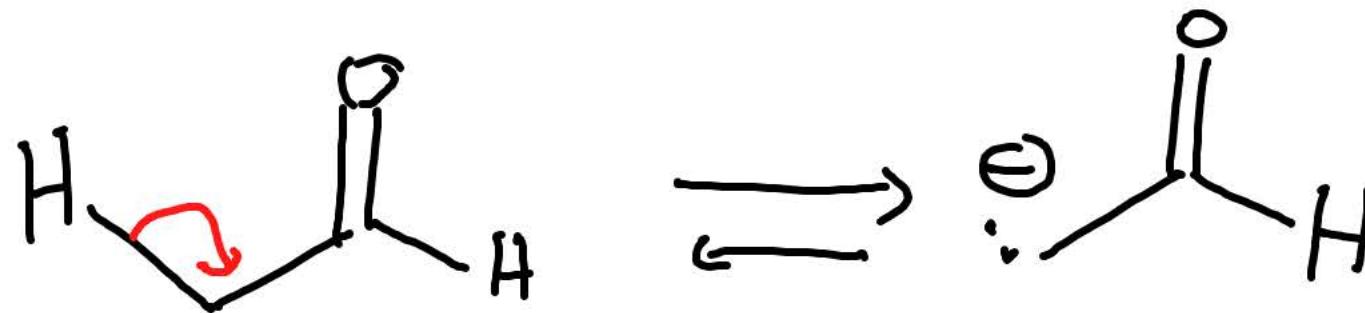


pKa = 11

EN(O) > EN(S), O stabilise mieux la charge moins, l'éthanol est plus acide

La molécule avec le soufre est en fait 100'000 fois plus acide!

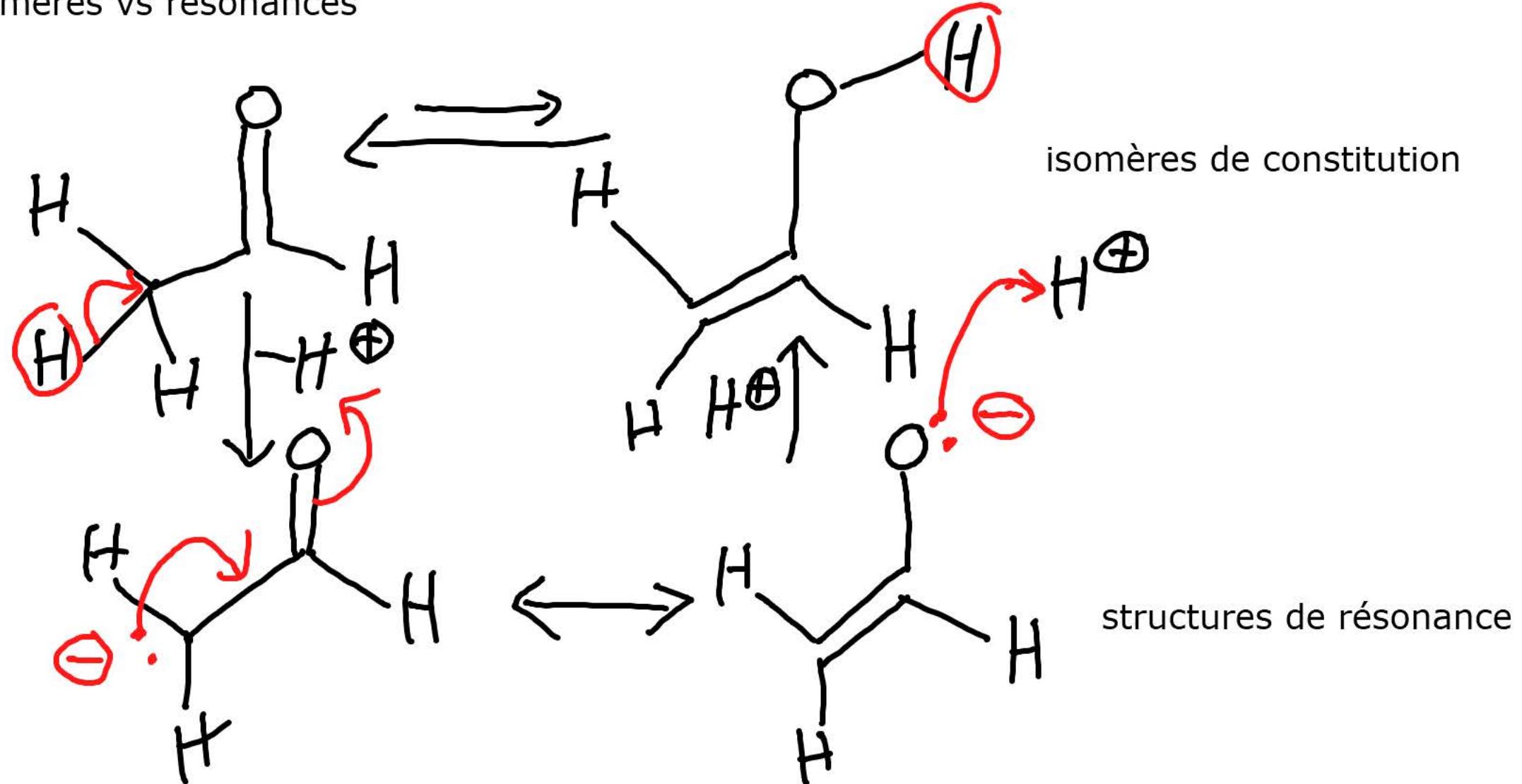
2ème cas



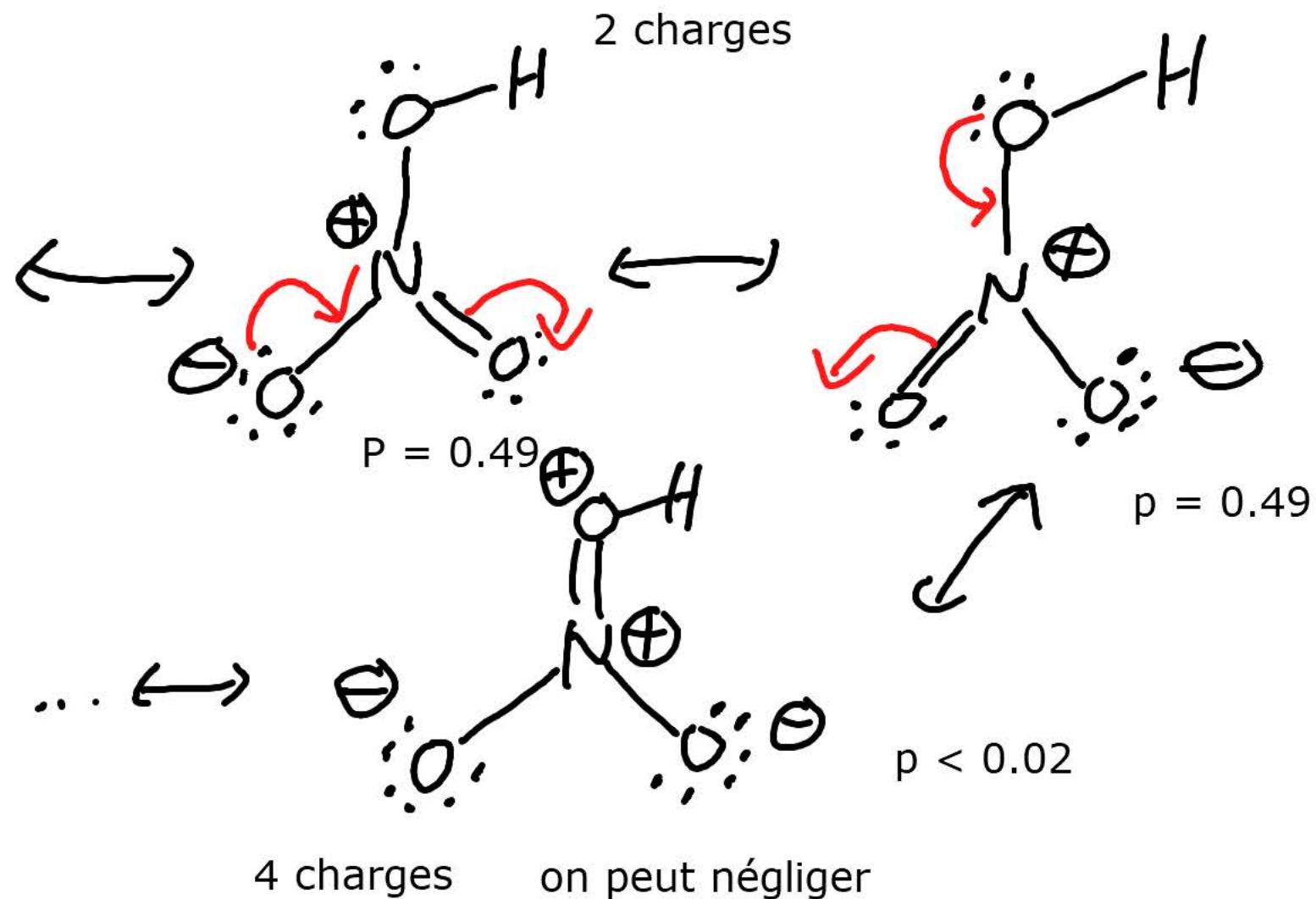
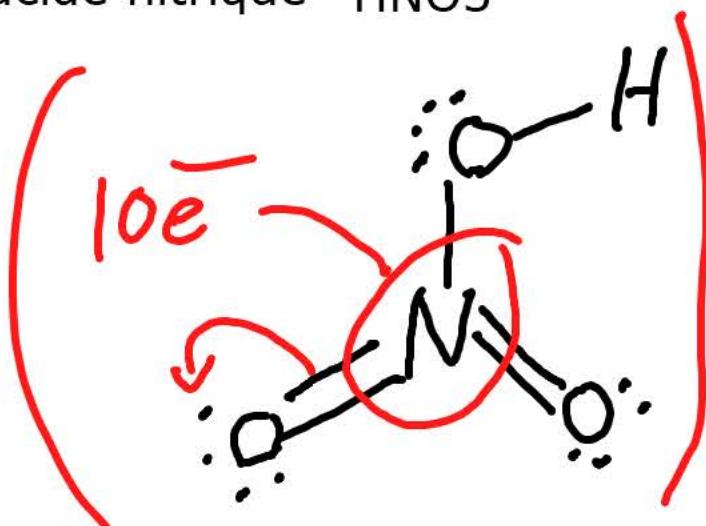
liaison C-H: mauvais acide (méthane: $pK_A = 45$)

mesure ici: $pK_A = 15$, 10 puissance 30 plus acide que méthane!

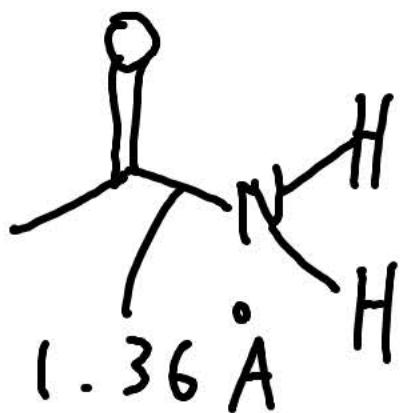
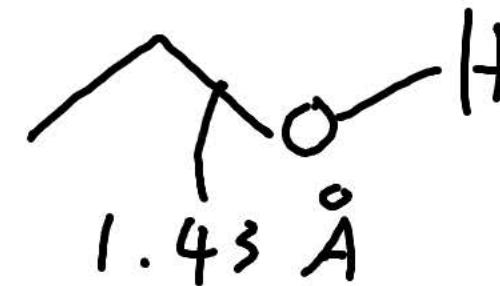
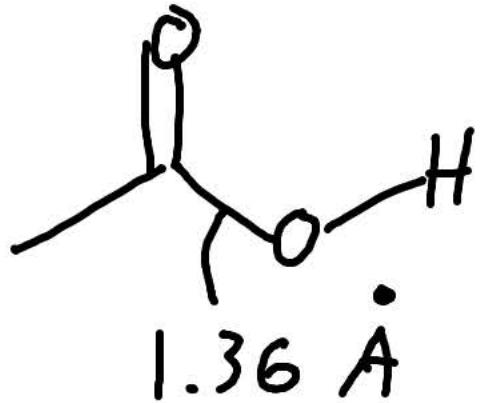
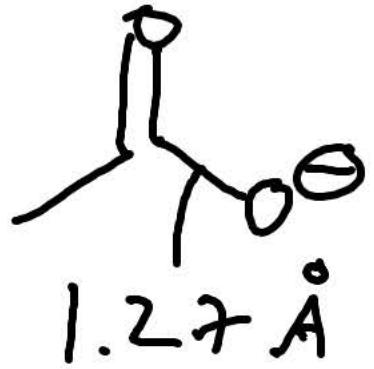
isomères vs résonances

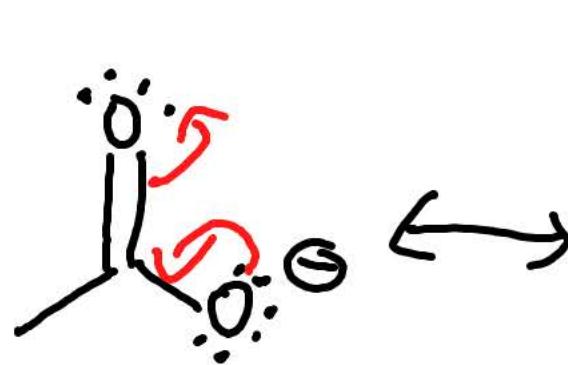


acide nitrique HNO_3

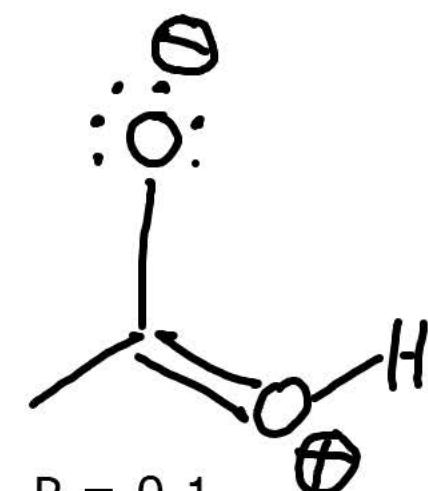
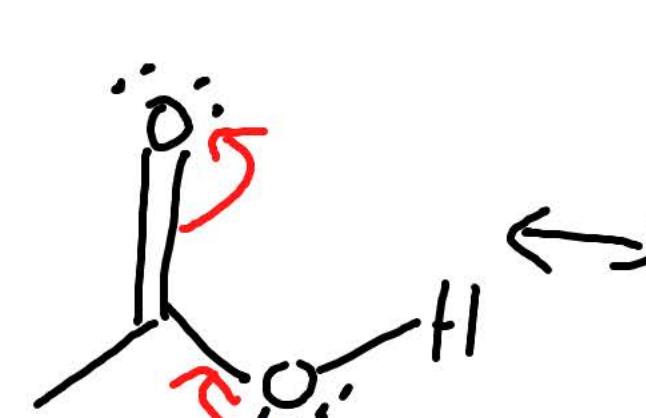


longueur de liaisons

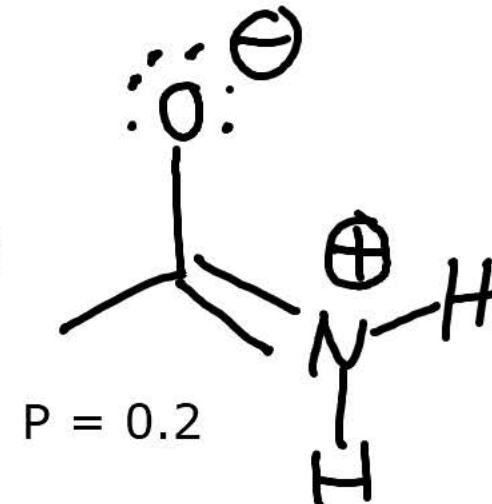
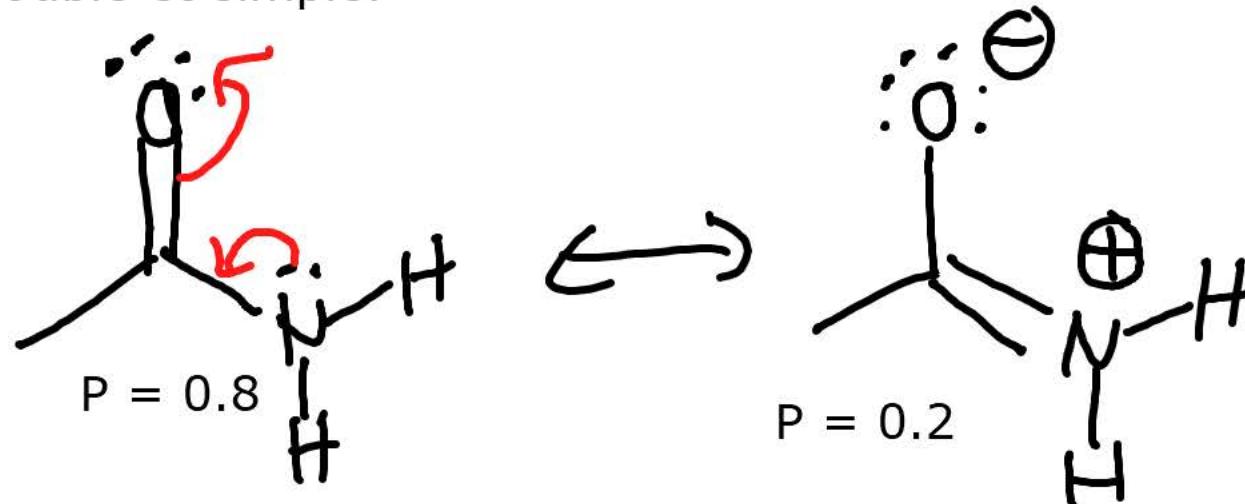




liaison intermédiaire entre
double et simple!

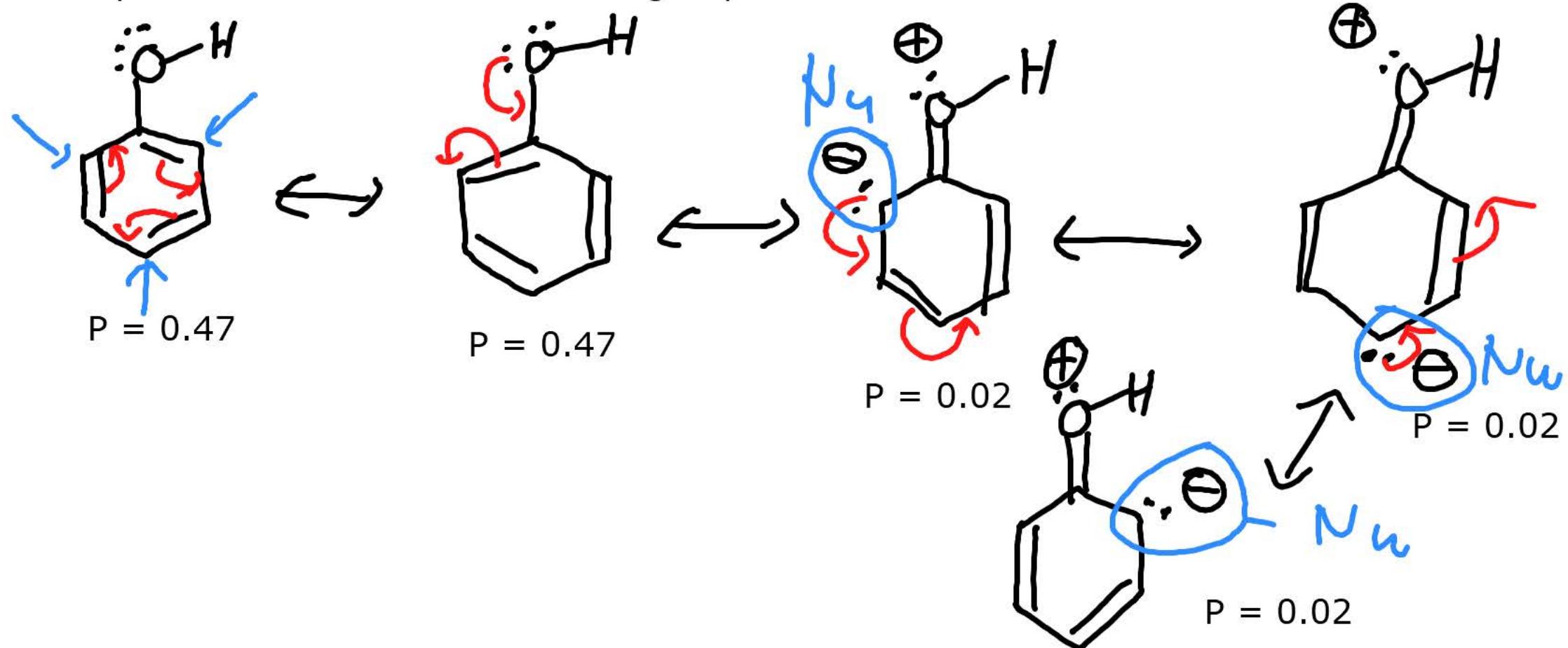


- sur O, + sur O



- sur O, + sur N
structure de résonance plus
stable

cas particulier: le benzène avec un groupe donneur d'électron



benzène avec groupe attracteur

