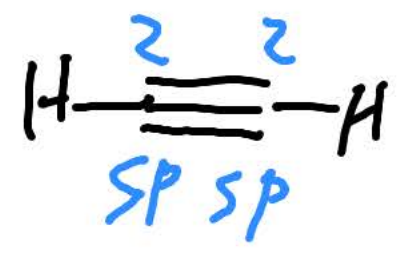
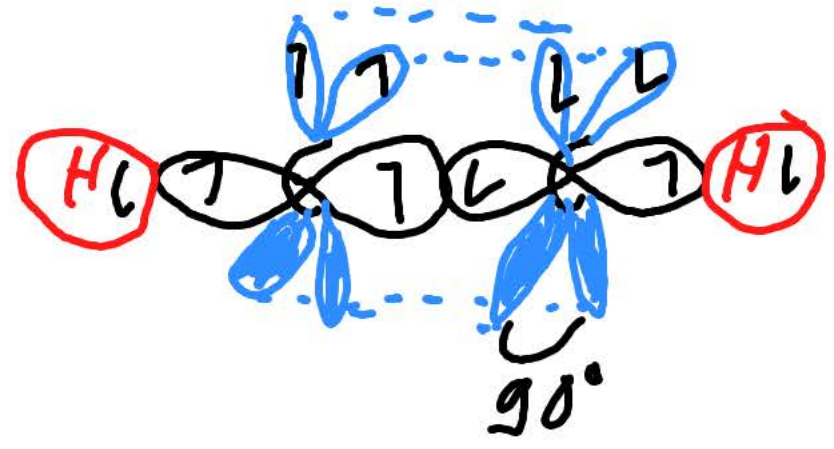


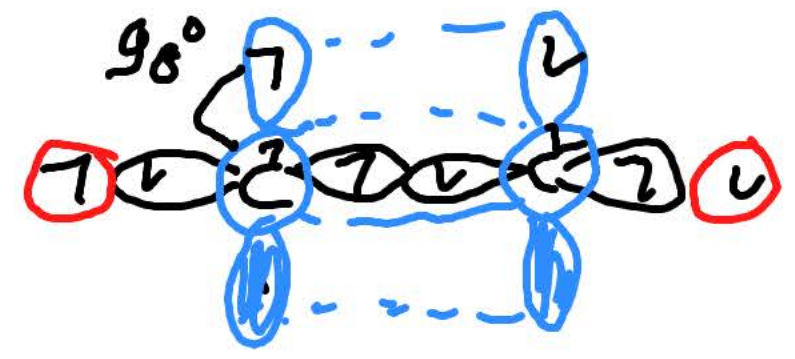
acétylène

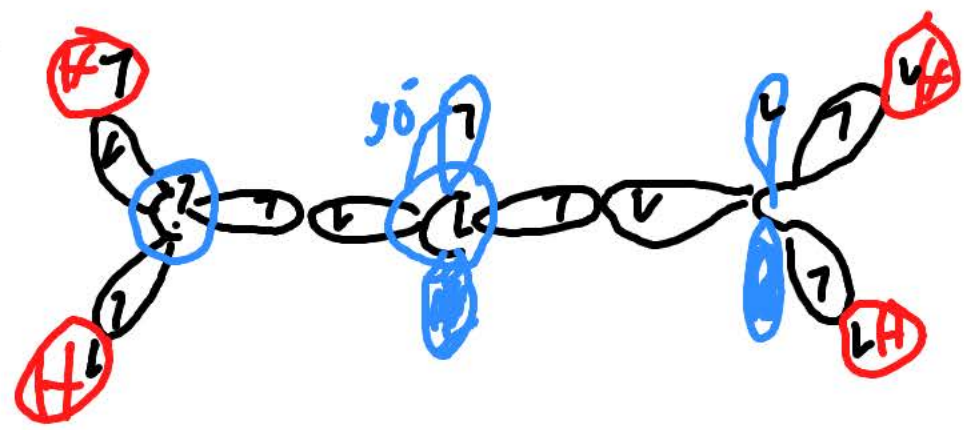
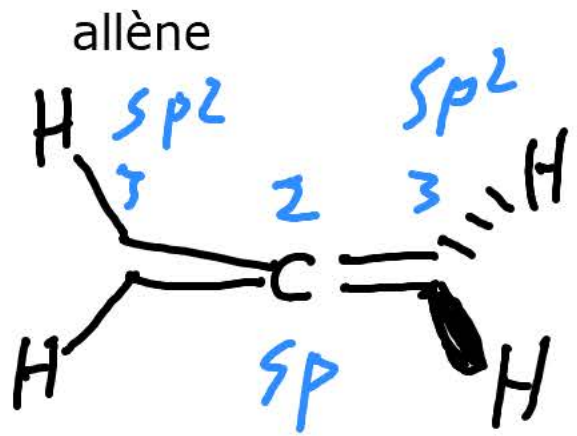


σ
 π



\equiv





4 étapes analyses des orbitales

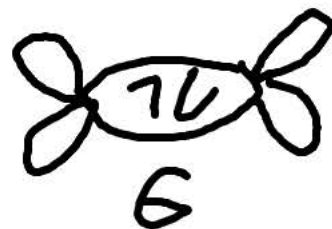
1) atome



2) interactions orbitales

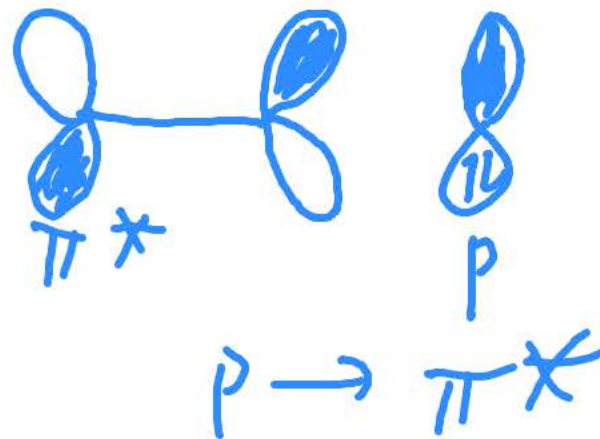


3) orbitales moléculaires



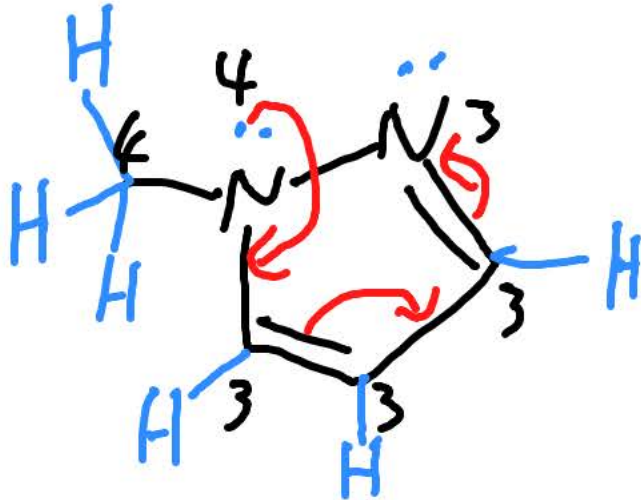
70-120 kcal/mol
liaisons

4) interactions orbitales secondaires



1-30 kcal/mol
résonances

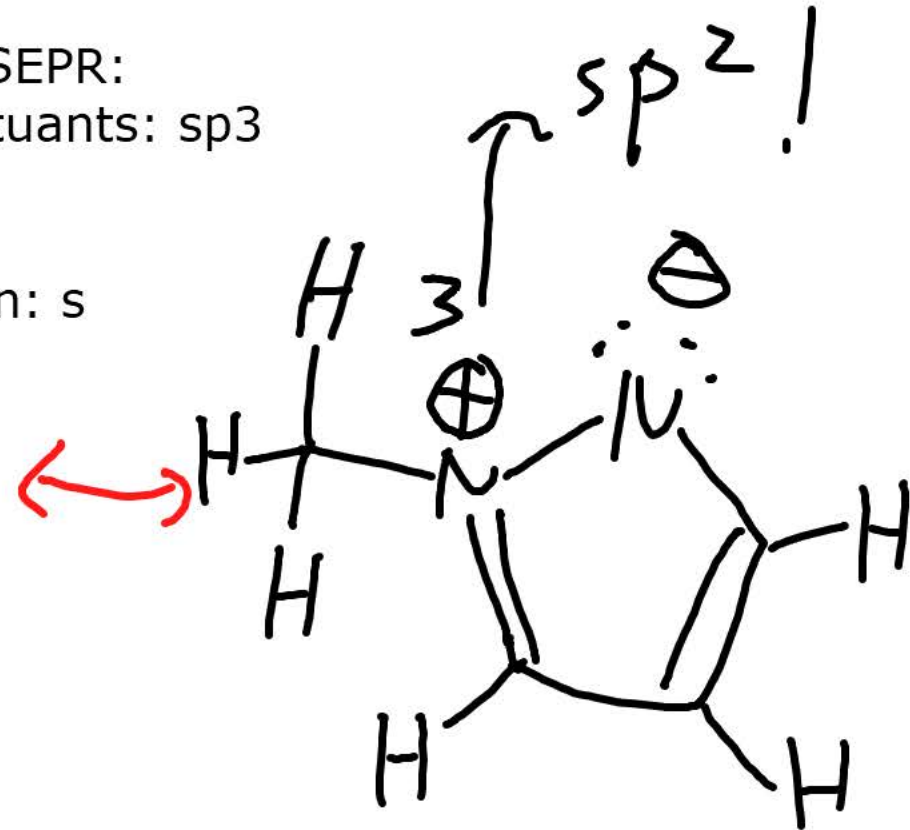
Quel est l'hybridisations des atomes?



1) ajouter les H et les paires!

attention: exceptions: changement de sp^3 en sp^2 pour permettre les résonances

Selon VSEPR:
4 substituants: sp^3
3: sp^2
2: sp
hydrogen: s



ici une exception.

dessiner les interactions orbitales

