

Dosage Example (3)

- 8) Molar mass of $\text{FeCl}_3 = 162.2 \text{ g/mol}$
- 9) Mass of FeCl_3 per hour $= 145 \times 162.2 = 23'519 \text{ g/h} = 23.5 \text{ kg/h}$
- 10) A solution of 40% FeCl_3 (w/v) is used $\rightarrow 23.5/0.4 = 58.75 \text{ L solution/h}$
- 11) That represents a volume of 1400 L/d or $\sim 1.4 \text{ m}^3/\text{day}$