

$$\rho_0 = \text{Mass of Rhodamine-WT} / \text{Unit Volume}$$

$$\frac{\text{Mass of Rhodamine-WT}}{\text{Unit Volume}} = \frac{\text{Volume of Rhodamine-WT}}{\underbrace{\text{Unit Volume}}_{0.2}} \cdot \text{Specific gravity} \cdot \rho_{\text{H}_2\text{O}}$$

$$= 0.2 \cdot 1.15 \text{ g/l} = 230 \text{ g/l} //$$