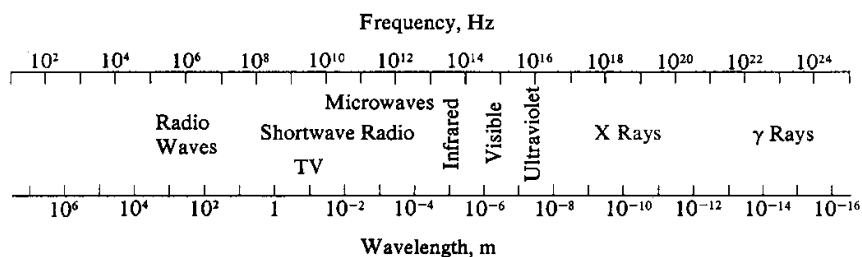


Quantum Chemistry

Exercises 1

- Radiation in the ultraviolet region of the electromagnetic spectrum is usually described in terms of wavelength, λ , and is given in units of angstrom ($1 \text{ \AA} = 10^{-10} \text{ m}$) or nanometers (10^{-9} m). Calculate ν , $\tilde{\nu}$, and the energy, E , for ultraviolet radiation with $\lambda=2000 \text{ \AA}$ and compare your results to the figure below.



- Chromium has a work function of 4.4 eV.
 - Calculate the maximum wavelength of light able to remove electron from a chromium surface.
 - Calculate the kinetic energy of electrons emitted from a chromium surface when it is irradiated with ultraviolet radiation of wavelength 2000 $\text{\AA}$.
 - What is the stopping potential for these electrons?
- A homogeneous light beam of wavelength $\lambda=300 \text{ nm}$ and intensity $5 \cdot 10^{-2} \text{ W m}^{-2}$ falls on a sodium surface. The photoelectric work function for sodium is 2.3 eV. Calculate:
 - the average number of photons striking the surface per m^2 and per second
 - the average number of electrons emitted per m^2 and per second
 - the maximum kinetic energy of the photoelectrons
- According to Planck's law the energy per unit time radiated per unit area *per unit frequency* by a black body at temperature T is given by:

$$I(\nu, T) = \frac{2h\nu^3}{c^2} \frac{1}{e^{\frac{h\nu}{kT}} - 1}$$

- From this, derive Wien's displacement law which relates the wavelength at which the intensity *per unit wavelength* of the radiation produced by a black body is at a maximum to the temperature of the black body.
 - Assuming that the sun can be described as a black body and knowing that the wavelength of maximum emission intensity is 500 nm, calculate the temperature of the sun.
- Show that the Lyman series occurs between 912 $\text{\AA}$ and 1216 $\text{\AA}$, that the Balmer series occurs between 3630 $\text{\AA}$ and 6563 $\text{\AA}$, and that the Paschen series occurs between 8210 $\text{\AA}$ and 18760 $\text{\AA}$. Identify the spectral regions to which these wavelengths correspond.

6. Calculate the de Broglie wavelength for:
- an electron with a kinetic energy of 100 eV
 - a proton with a kinetic energy of 10^5 eV