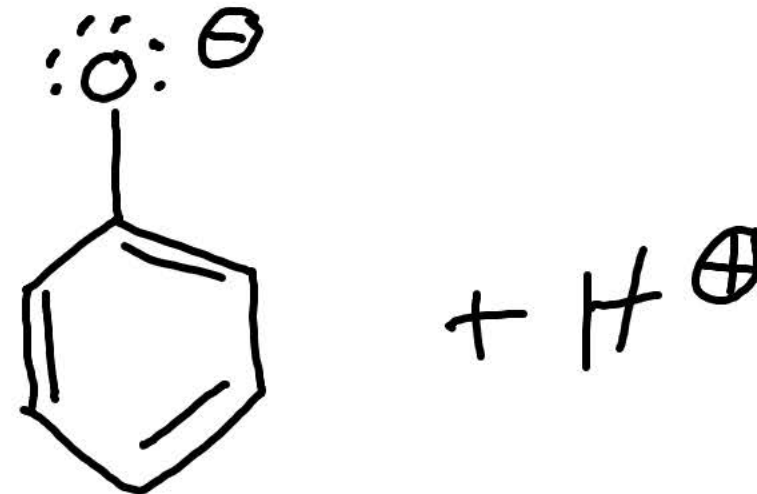
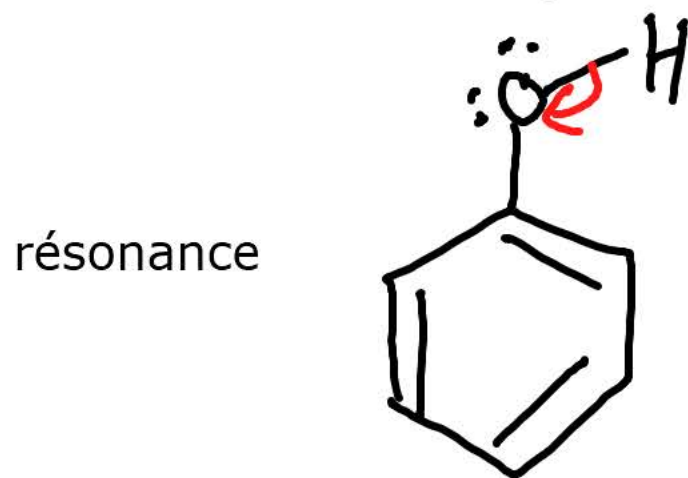
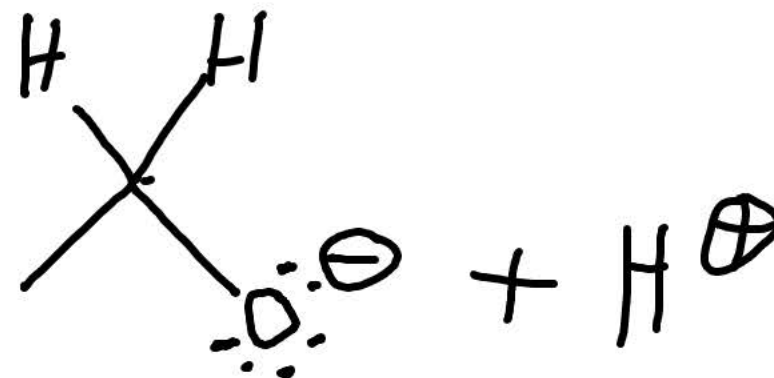
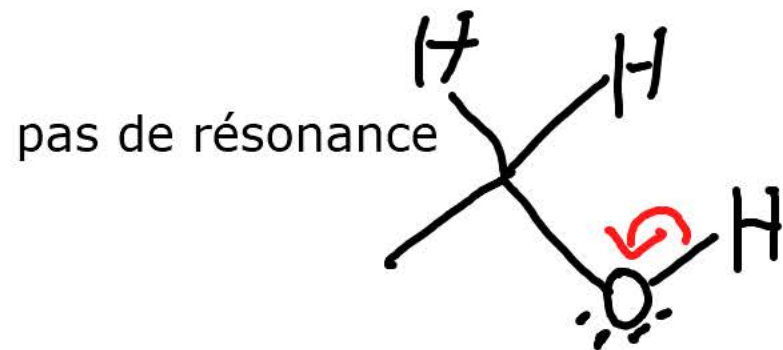
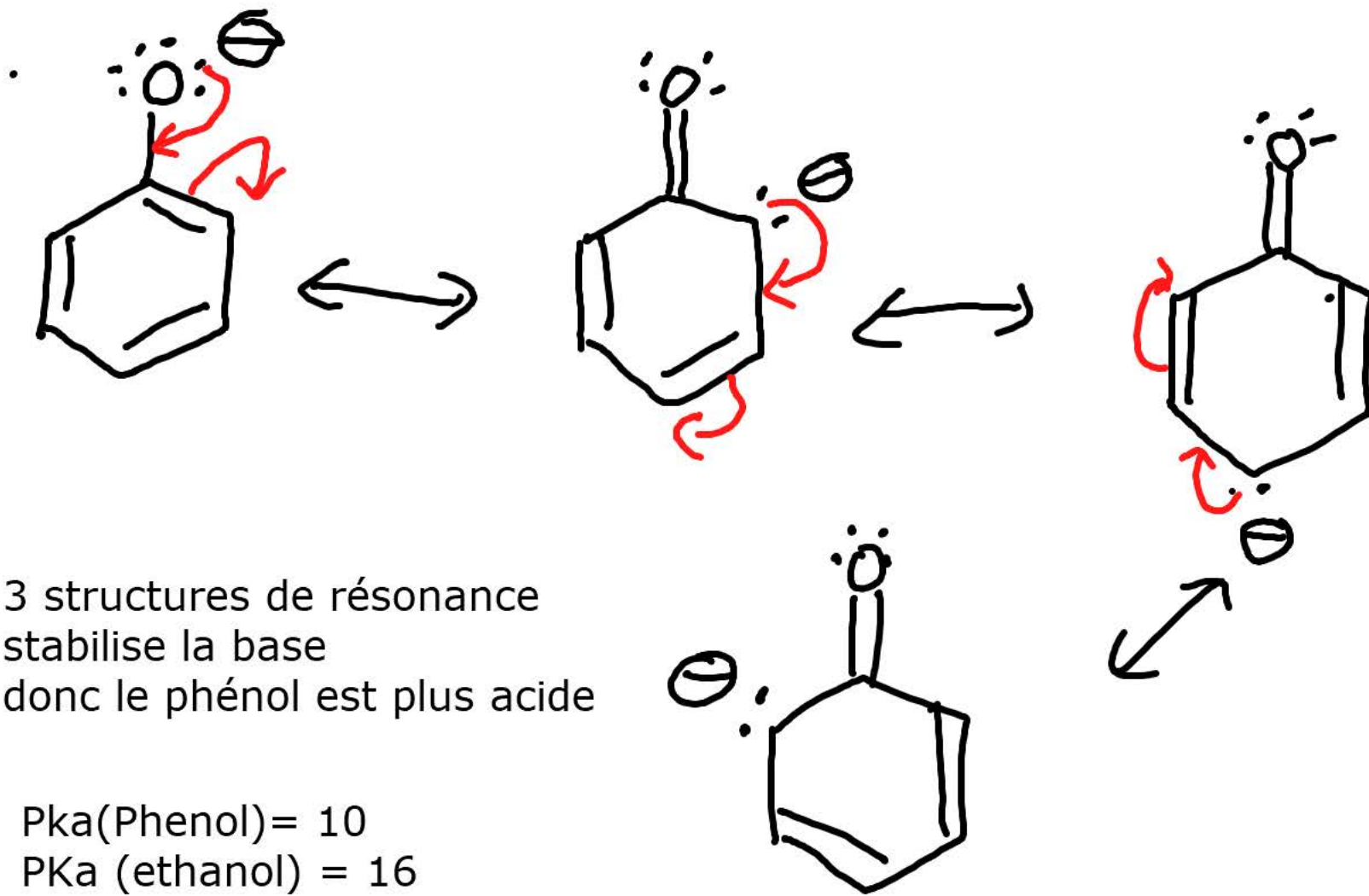


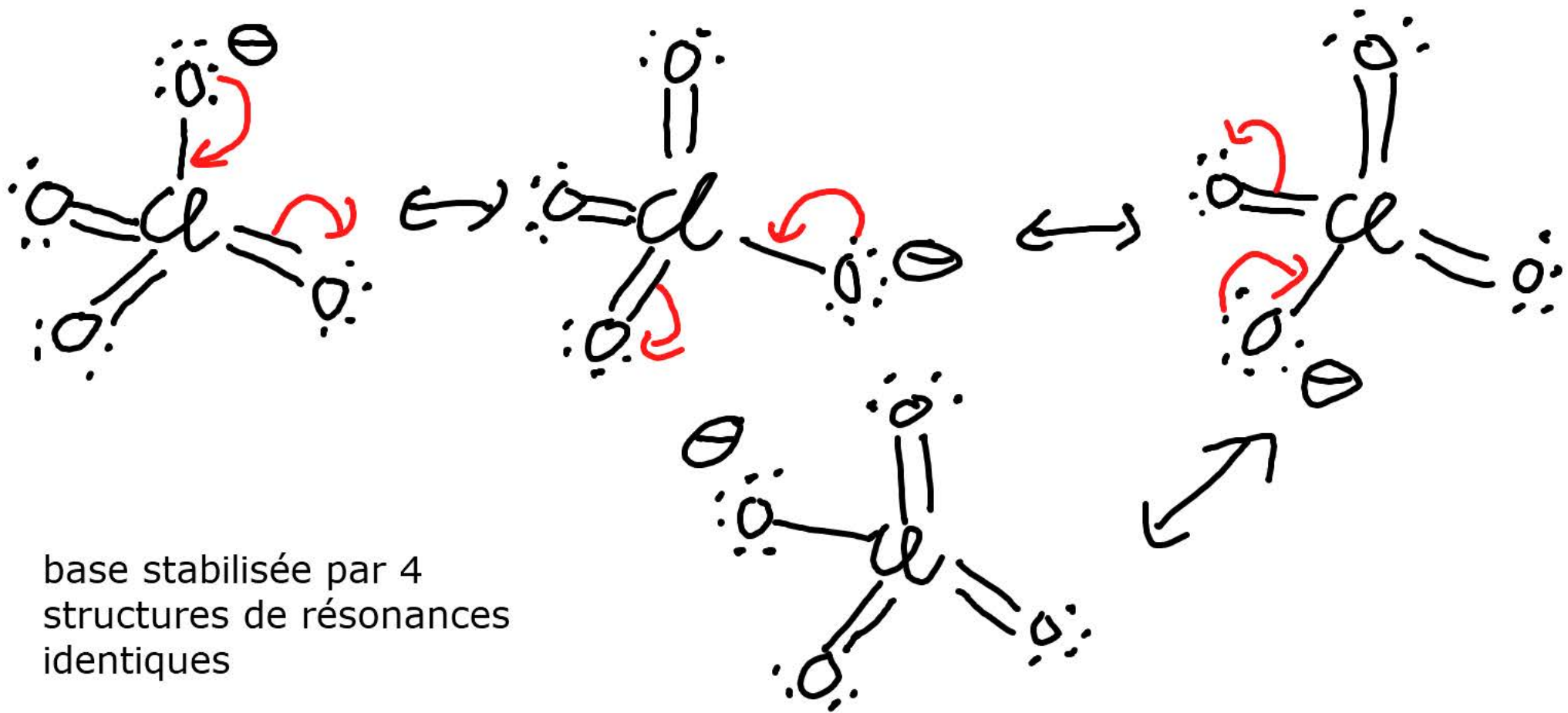
éthanol vs phénol, qui est plus acide



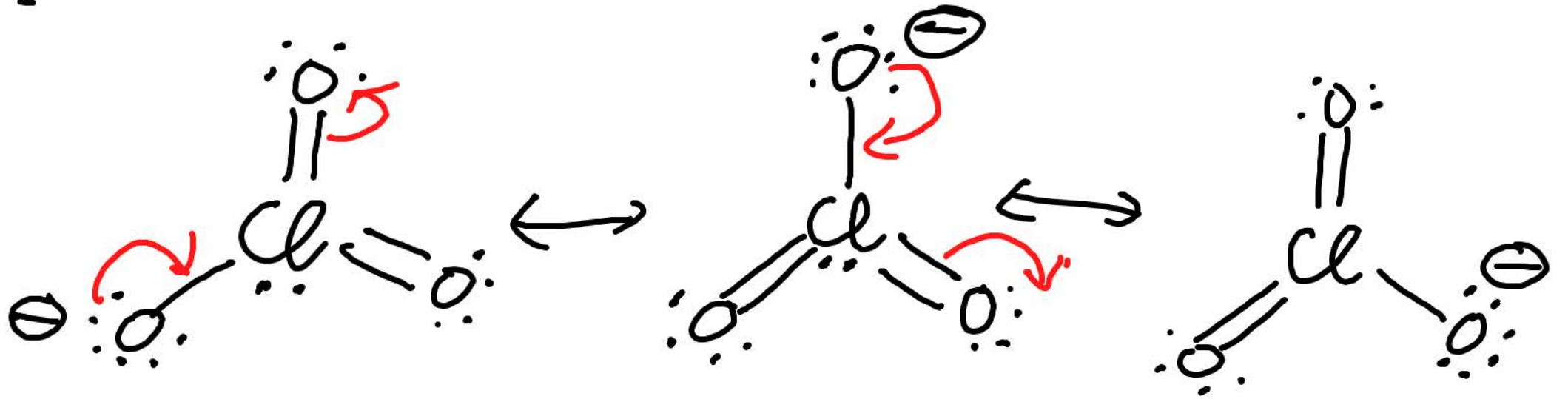
pas de différence d'électronégativité



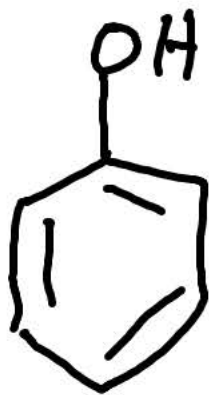
ClO₄⁻ ou ClO₃⁻: qui est plus basique?



base stabilisée par 4
structures de résonances
identiques



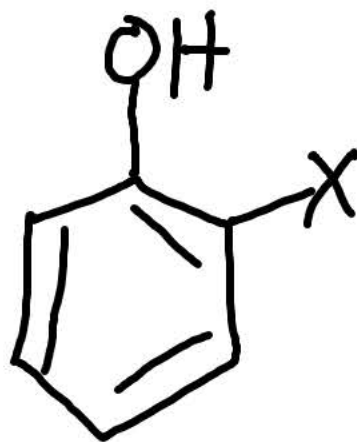
3 structures de résonance
moins stable que ClO_4^- , donc
base plus forte



$\text{PKa} = 10$

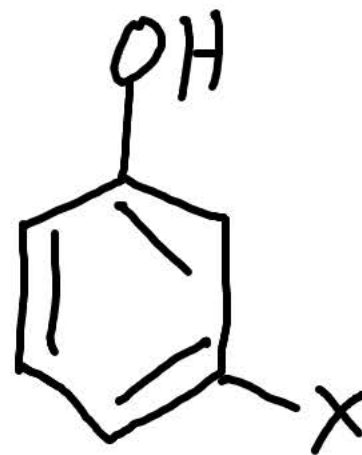
$\text{X} = \text{F}$

$\text{X} = \text{NO}_2$



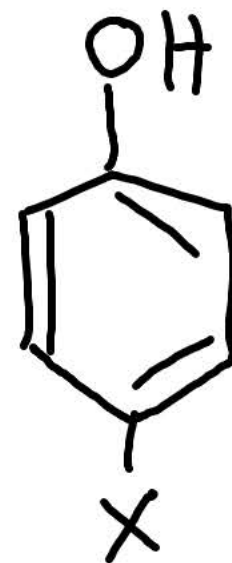
$\text{PKa} = 8.7$

$\text{PKa} = 7.2$



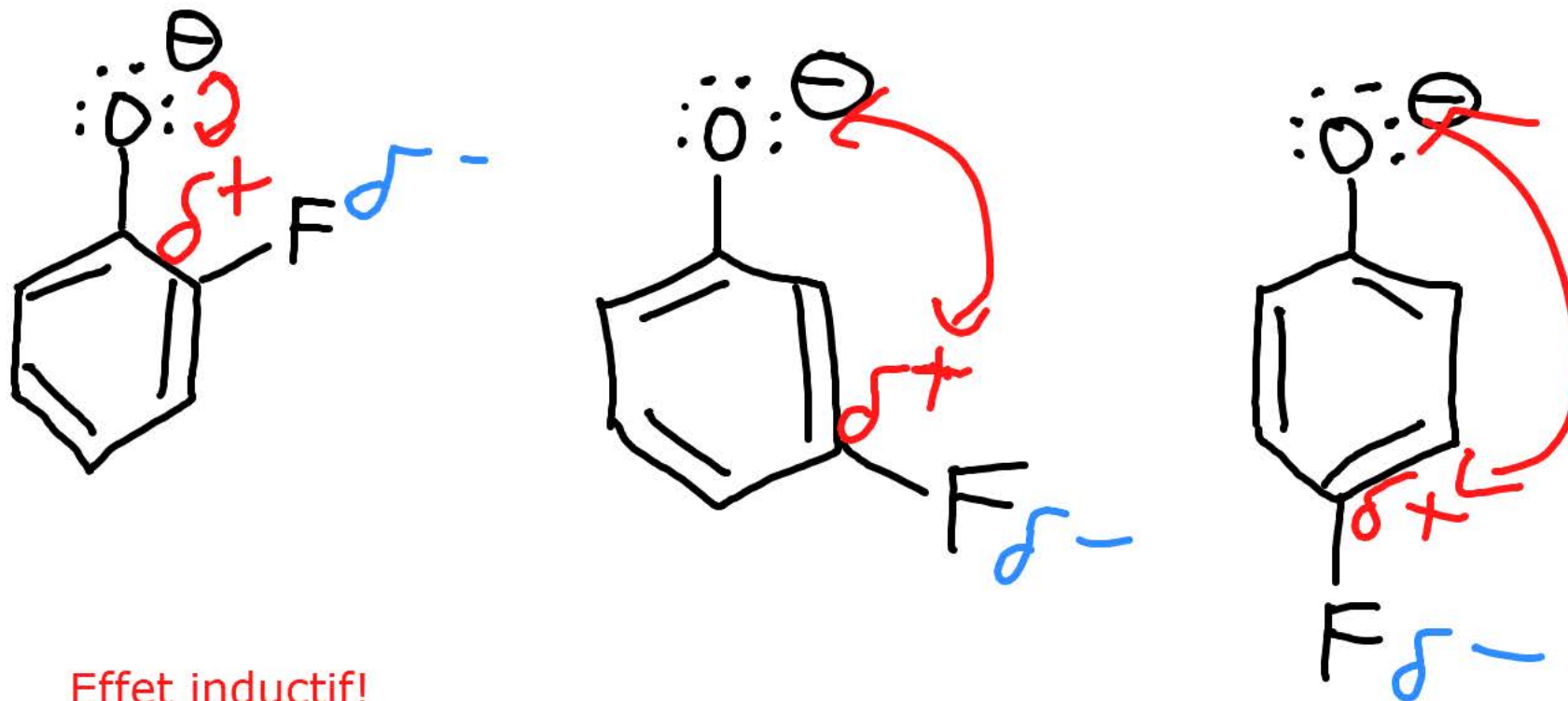
$\text{PKa} = 9.3$

$\text{PKa} = 8.4$



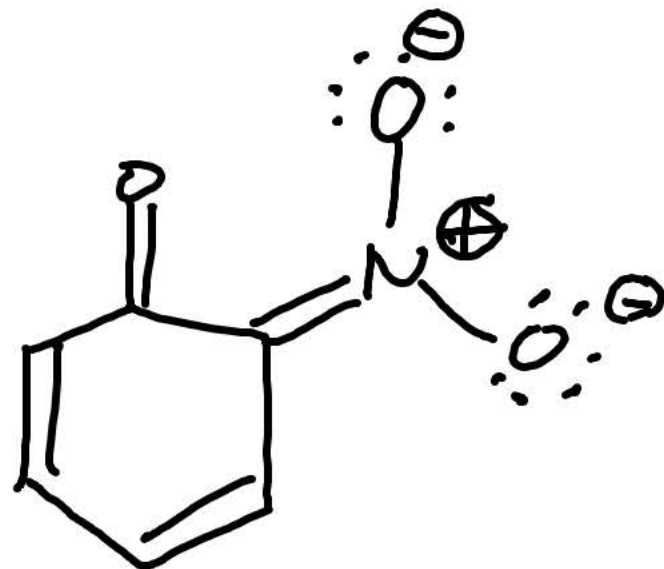
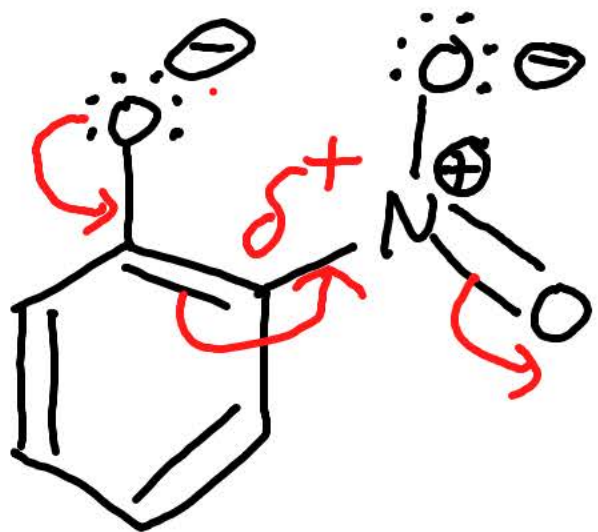
$\text{PKa} = 9.7$

$\text{PKa} = 7.2$

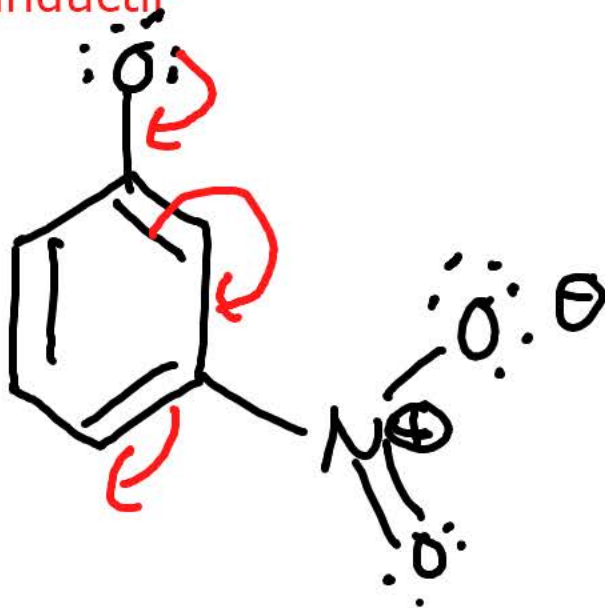


Effet inductif!

Effet stabilisant sur la base, qui diminue avec la distance
Quand la base est plus stable, l'acide est plus fort, le
P_kA est plus bas

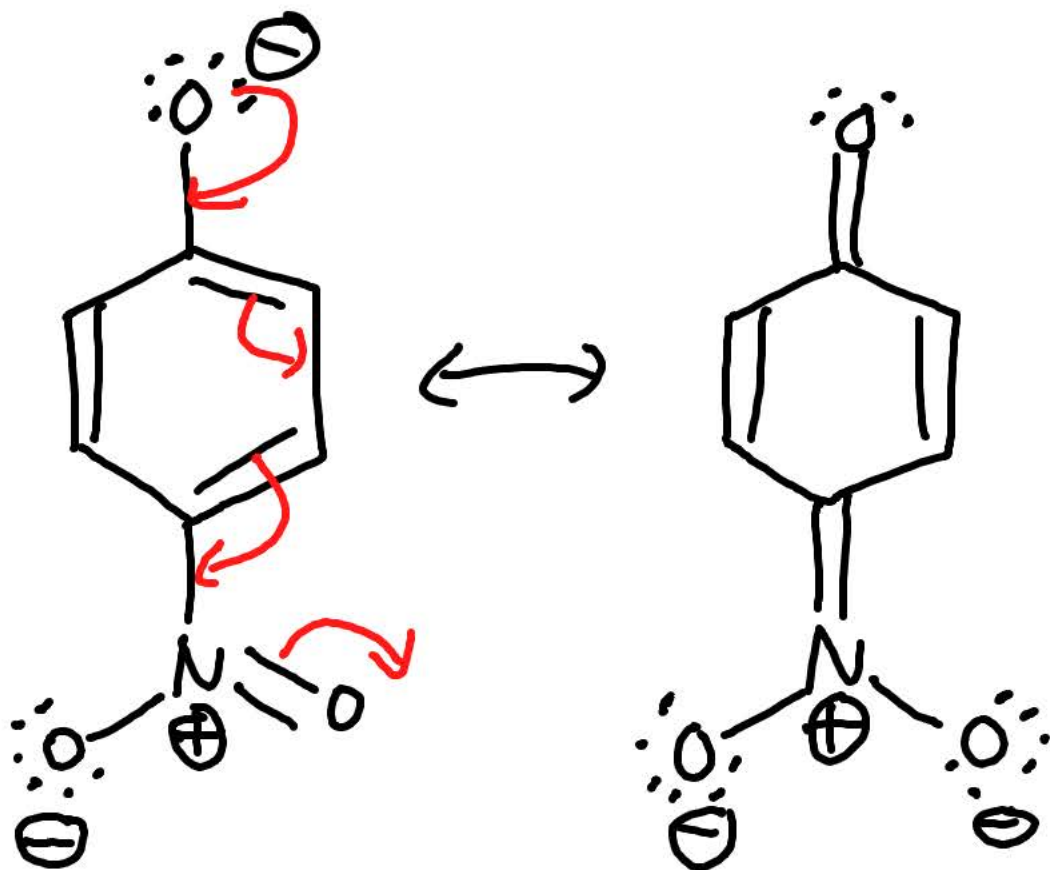


effet inductif



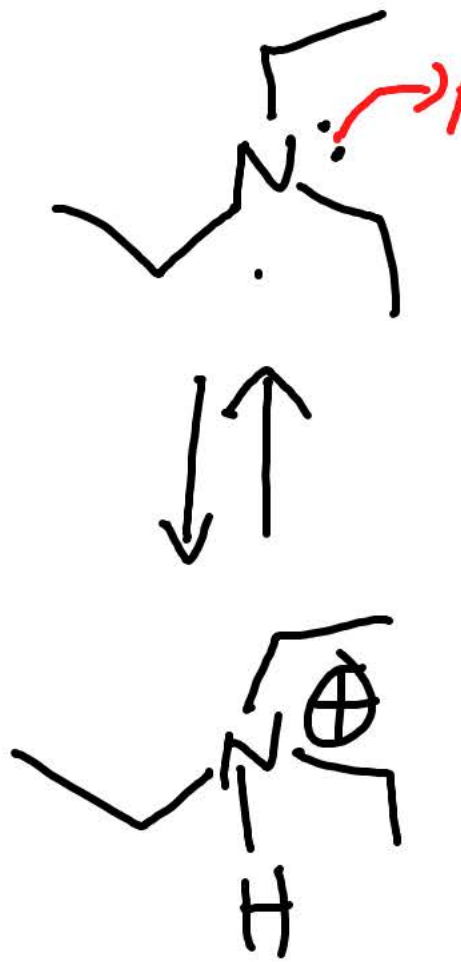
résonance en plus, bonne
résonance avec moins sur O
base stabilisée, acide plus fort

pas de résonance sur le nitro!
une structure de résonance en
moins, base moins stabilisée,
base est moins forte



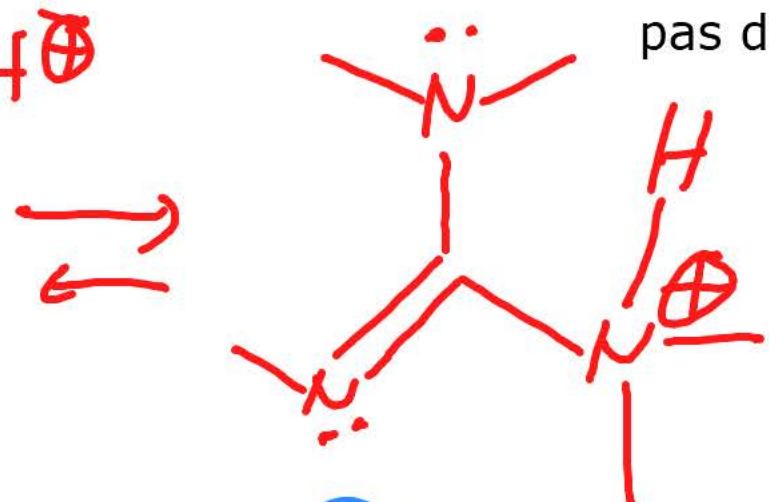
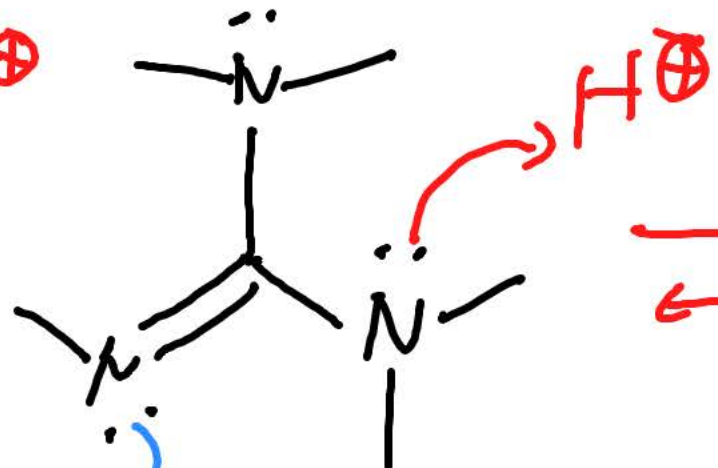
Très bonne résonance avec le
nitro
base est stabilisée
et l'acide plus fort

$\text{PkaH} = 11$

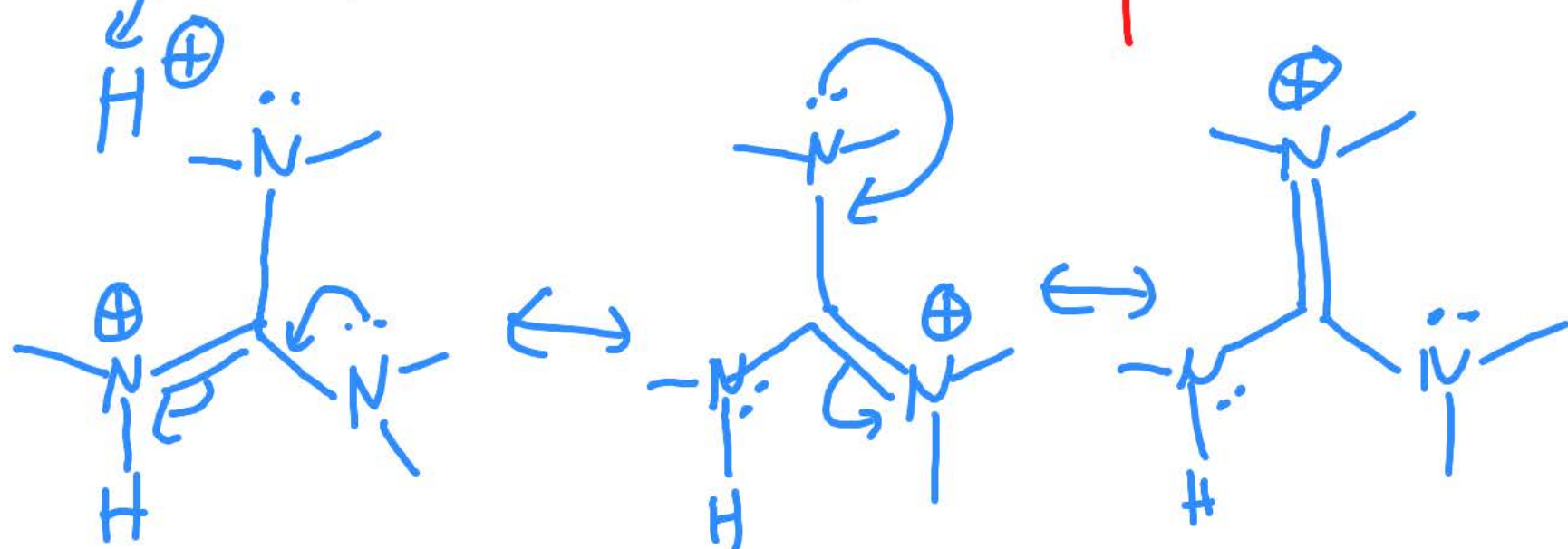


pas de résonance

$\text{PkaH} = 14$



pas de résonance



très bonnes résonances, l'acide est plus stable, donc la base plus forte